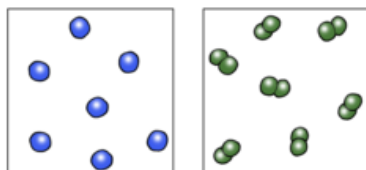


Atoms, elements and compounds

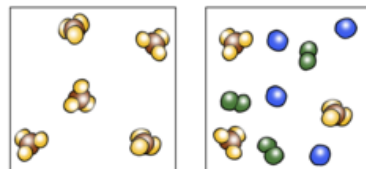
All substances are made of **atoms** that cannot be chemically broken down. It is the smallest part of an **element**.

Elements are made of only one type of atom. Each element has its own **symbol**.
e.g. Na is sodium.

Compounds contain more than one type of atom.
Naming compounds-
Two elements = **ide**
e.g. Na₂S Sodium sulphide
Two or more including oxygen = **ate**
e.g. Na₂SO₄ = sodium sulphate



a) Atoms of an element b) Molecules of an element



c) Molecules of a compound d) Mixture of elements and a compound



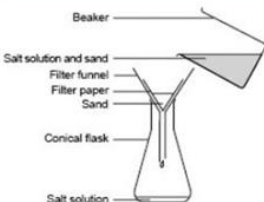
Small numbers (subscripts) after symbols tell you how many of the element BEFORE the number.

Separating mixtures

A mixture consists of **two or more** elements or compounds **not** chemically combined together.

Filtration

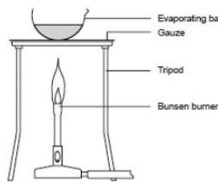
This technique separates substances that are insoluble in a solvent from those that are soluble



Example - filtering a mixture of sand, salt and water to collect the sand

Crystallisation

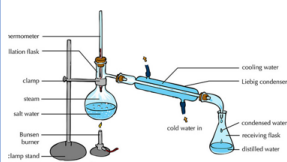
This technique separates a soluble substance from a solvent by heating



Example - crystallisation of sodium chloride from salt solution

Simple distillation

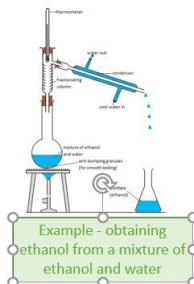
This technique separates a liquid from a mixture by evaporation followed by condensation



Example - obtaining water from sea water

Fractional distillation

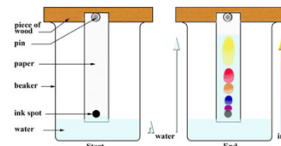
This technique differs from distillation only in that it separates a mixture into a number of different parts, called fractions.



Example - obtaining ethanol from a mixture of ethanol and water

Chromatography

This technique separates small amounts of dissolved substances by running a solvent along absorbent paper



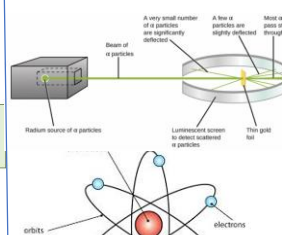
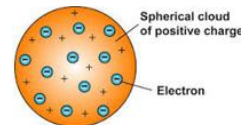
Example - separating the different colours in ink

C1 Atomic Structure

Development of Atomic Model

Dalton – atoms can't be divided

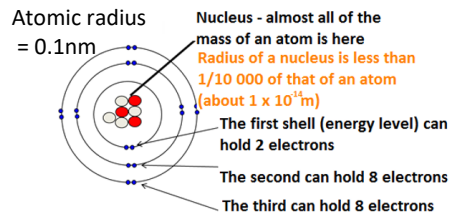
JJ Thompson discovered electrons – Plum pudding model



Geiger-Marsden The Nuclear Model of the Atom

Bohr – electrons in shells

Chadwick – the neutron



Subatomic Particles

	Mass	Charge	Location
Proton	1	+	nucleus
Neutron	1	0	nucleus
Electron	Very small	-	shells

Mass number = Number of protons and neutrons → ⁷Li
Atomic number = Number of protons → ₃

Number of protons(+) = Number of electrons (-)

Number of neutrons = mass number – atomic number

⁷Li
3

Protons = 3
Electrons = 3
Neutrons = 4

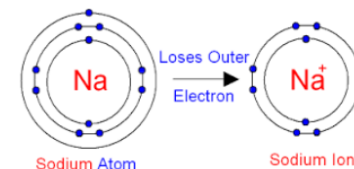
Isotopes

Different mass numbers

Same atomic number

¹²6C ¹³6C ¹⁴6C

Atoms lose or gain electrons to form ions



1nm = 1x10⁻⁹m